



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

CANDIDATE  
NAME

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NUMBER

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**COMBINED SCIENCE**

**0653/02**

Paper 2 (Core)

**May/June 2009**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

No Additional Materials are required.



**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

**DO NOT WRITE IN ANY BARCODES.**

Answer **all** questions.

A copy of the Periodic Table is printed on page 24.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
1	
2	
3	
4	
5	
6	
7	
8	
9	
<b>Total</b>	

This document consists of **21** printed pages and **3** blank pages.



1 Fig. 1.1 shows a section through a tooth.

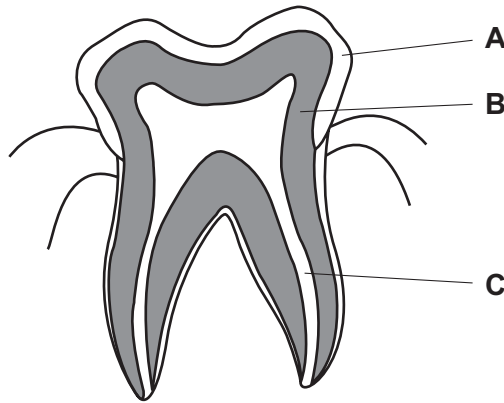


Fig. 1.1

(a) Name parts **A**, **B** and **C**.

**A** .....

**B** .....

**C** .....

[3]

(b) Explain how teeth help with digestion.

.....

.....

..... [2]

(c) Name **one** mineral and **one** vitamin that are essential for the growth of strong teeth and bones.

mineral .....

vitamin .....

[2]

- 2 (a) A student investigated how a change in potential difference across a lamp affected the current flowing through it.

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She used wires to connect the components shown in Fig. 2.1 to make a circuit.

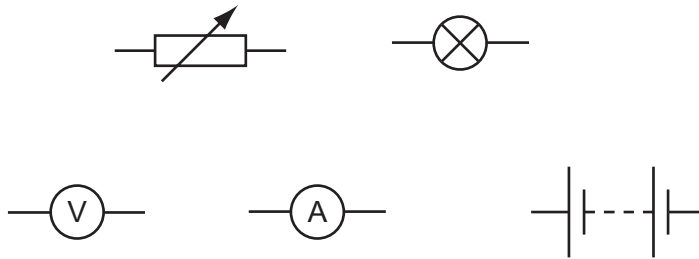


Fig. 2.1

- (i) Using the correct symbols from Fig. 2.1, draw a diagram to show the circuit she used.

[3]

- (ii) Explain why the variable resistor is included in the circuit.

.....  
 ..... [1]

(iii) Her results are shown in Table 2.1.

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**Table 2.1**

potential difference across lamp/V	current through lamp/A	resistance of lamp filament/ $\Omega$
4	1.2	3.3
8	1.5	
12	1.7	7.1

Complete the table by calculating the missing resistance and writing your answer in the empty box.

State the formula that you use and show your working.

formula

working

[2]

(b) Electricity can kill.

Identify and explain the electrical hazard shown in Fig. 2.2.

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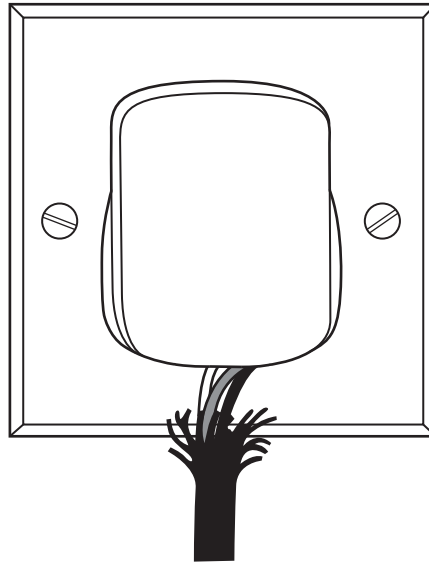


Fig. 2.2

.....

.....

.....

..... [2]

3 (a) The names of six elements are shown below.

**carbon      chlorine      cobalt      neon      silicon      sodium**

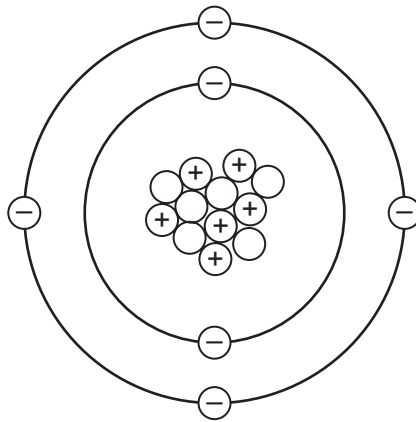
Choose the element from the list

which is the least reactive, .....

which is used to sterilise drinking water, .....

which is a metal that forms coloured compounds. .... [3]

(b) Fig. 3.1 shows a diagram of an atom.



**Fig. 3.1**

(i) State the nucleon number (mass number) of the atom shown in Fig. 3.1.  
..... [1]

(ii) State the name of the element made of atoms like the one in Fig. 3.1.  
Explain your answer briefly.  
element .....

explanation .....

..... [2]

(c) Fig. 3.2 shows a test for a gas which is produced when a solid element **A** reacts in a solution **B**.

For  
Examiner's  
Use

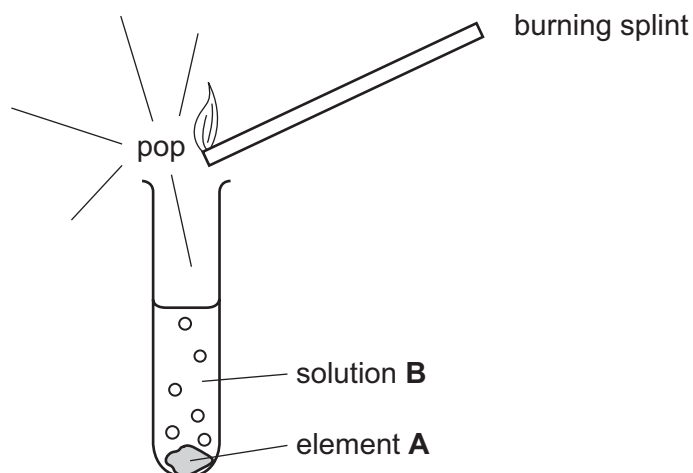


Fig. 3.2

Name the gas produced in this reaction, and suggest the names of element **A** and solution **B**.

gas .....

element **A** .....

solution **B** .....

[3]

4 Fig. 4.1 shows an arum lily.

Arum lilies have flowers that are pollinated by insects.

There are many tiny flowers on a stalk, inside a large white structure called a spathe.

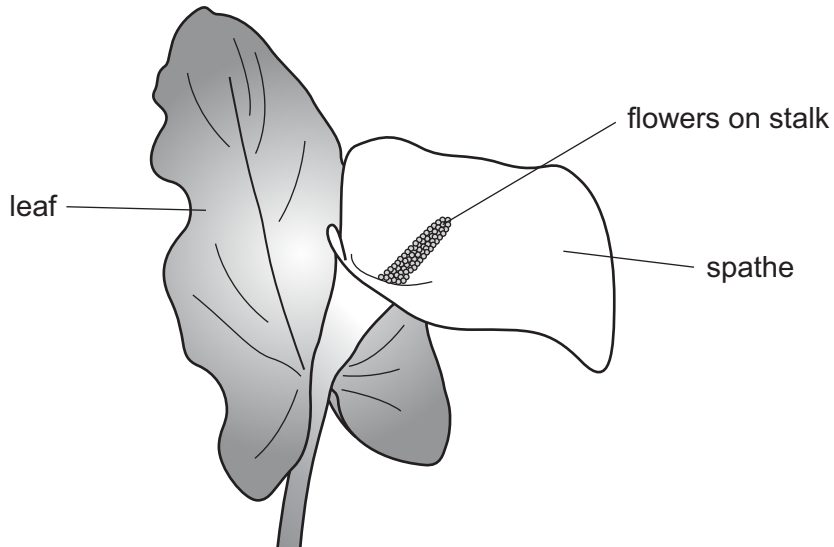


Fig. 4.1

(a) (i) Name the part of the flower in which pollen is made.

..... [1]

(ii) What does a pollen grain contain?

..... [1]

(iii) Explain the meaning of the term *pollination*.

.....  
 .....  
 ..... [2]



- (b) Arum lilies produce heat energy to raise the temperature of the flowers. This helps to attract insects to the flowers. They use respiration to do this.

For  
Examiner's  
Use

A researcher investigated whether there was a relationship between the temperature of the flowers inside an arum lily spathe and the rate of oxygen use.

He took 15 arum lilies, and measured the temperature and rate of oxygen use for each one.

Fig. 4.2 shows his results.

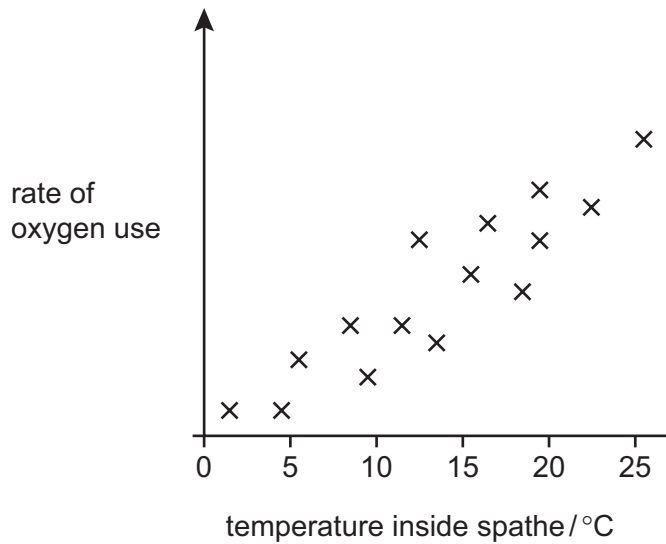


Fig. 4.2

- (i) Describe the relationship between the temperature inside the spathe and the rate of oxygen use by the arum lily.

.....  
..... [1]

- (ii) Explain the reasons for the relationship you have described.

.....  
.....  
.....  
..... [2]

(c) The fuel that the arum lilies use to produce the heat energy is glucose.

Describe how the lilies obtain a supply of glucose.

.....  
.....  
..... [2]

(d) The leaves of arum lilies contain palisade cells, which are typical plant cells containing chloroplasts.

Complete the diagram of a palisade cell. Include these structures in your labels.

**cell membrane**

**cell wall**

**chloroplast**

**cytoplasm**

**nucleus**

**vacuole**



[4]

**Please turn over for Question 5.**

5 The Vikings, who lived in Scandinavia about 1200 years ago, sailed in boats called longships across the sea to Britain.

(a) (i) They travelled 900 km in 150 hours.

Calculate their average speed for this journey.

State the formula that you use and show your working.

formula

working

..... km/h [2]

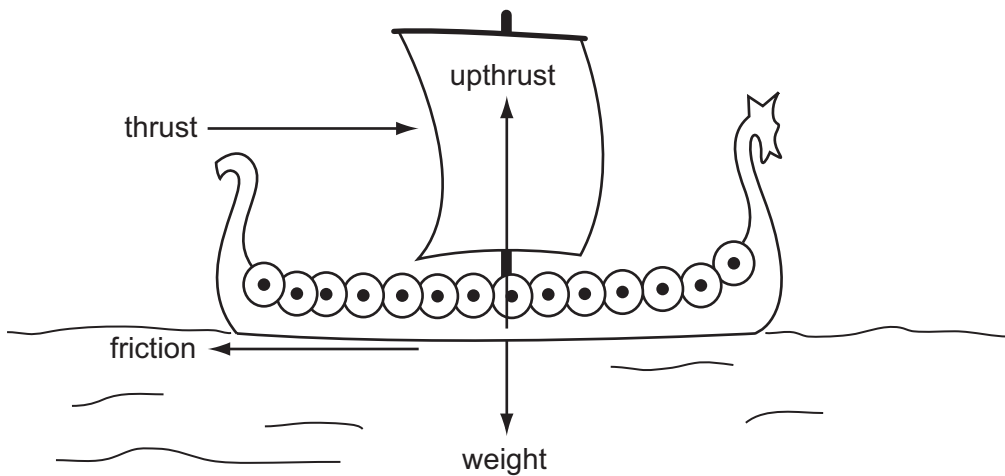
(ii) At one stage on their journey they were travelling at 7.2 km/h.

Calculate their speed in m/s.

Show your working.

..... m/s [1]

(b) A longship was moving at constant speed. The diagram shows four forces acting on it.



Name **two** forces which must be equal in size.

..... [1]

(c) The Vikings used animal furs to make clothes to keep them warm.

Explain in terms of conduction and convection how fur clothes would have kept the Vikings warm.

.....  
.....  
.....  
..... [3]

(d) The volume of the wood used to construct the longship was 9 m<sup>3</sup>.

If the density of the wood was 800 kg/m<sup>3</sup>, calculate the mass of the wood used.

State the formula that you use and show your working.

formula

working

..... kg [2]

(e) The major energy source used to propel the longship was the wind. Wind is a renewable energy source.

(i) Name **one** other renewable energy source.

..... [1]

(ii) Name **one** non-renewable energy source.

..... [1]

- 6 (a) Fig. 6.1 shows industrial apparatus used for the fractional distillation of petroleum (crude oil).

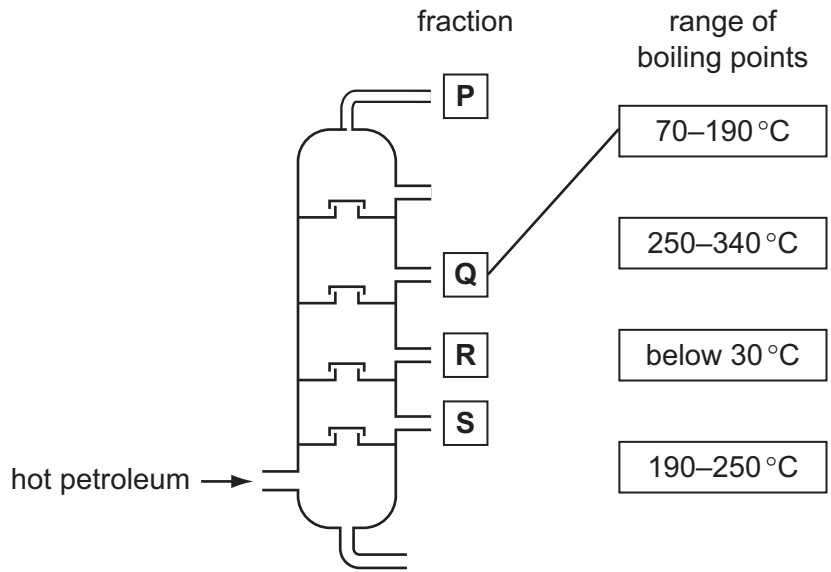
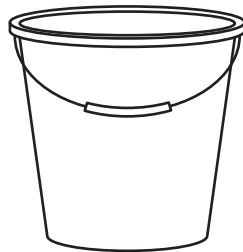


Fig. 6.1

Draw lines on Fig. 6.1 connecting the fractions, **P**, **Q**, **R** and **S** to the correct boiling point range. The line for fraction **Q** has been drawn for you. [2]

- (b) Plastics and steel are both used to make buckets.



bucket

Fig. 6.2

- (i) Suggest **one** reason why plastics are suitable materials from which to make buckets.

.....  
 ..... [1]

- (ii) Buckets made from steel must be protected from rusting.

Name the element and the compound which react with mild steel to form rust.

element .....

compound .....

[2]

- (iii) Describe briefly **one** suitable method of protecting a steel bucket from rusting.

.....

.....

[1]

- (iv) Name the element which is oxidised when rust forms.

.....

[1]

- (v) Name the alloy from which cutlery is made.



cutlery

**Fig. 6.3**

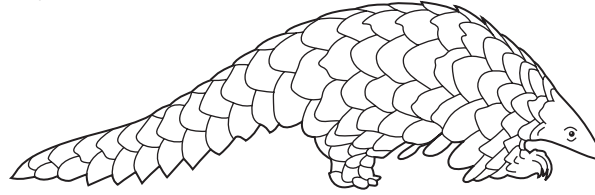
.....

[1]

7 Read the following description of a food web.

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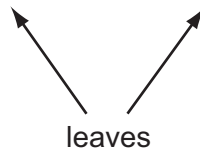
- Ants collect leaves from trees and take them into their nests.
- A fungus grows on the leaves and breaks them down.
- The ants eat the leaves, and also the fungus.
- Small birds eat the ants, and hawks eat the small birds.
- Pangolins eat only ants.



a pangolin

**Fig. 7.1**

(a) In the space below, complete a food web that includes all of the organisms described in Fig. 7.1.



[3]



(b) (i) Name the producer in this food web.

..... [1]

(ii) Name a decomposer in this food web.

..... [1]

(c) Pangolins are becoming rare in some parts of the world.

Use the information in Fig. 7.1, and your own knowledge, to explain why it is important to prevent deforestation if we want to conserve pangolins.

.....  
.....  
.....  
..... [2]

For  
Examiner's  
Use

- 8 (a) A hotel has a lift (elevator). It moves through a vertical height of 3 m between each floor.

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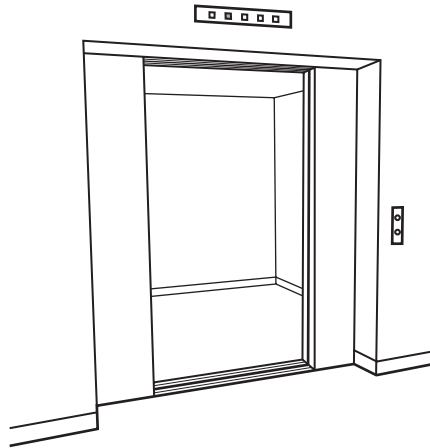


Fig. 8.1

- (i) A passenger travels in the lift. The passenger has a mass of 80 kg and weighs 800 N. The mass of the empty lift is 1200 kg.

Calculate the total weight of the passenger and lift.

Show your working.

..... N [2]

- (ii) Calculate the work done when the lift and passenger move up three floors, from Floor 1 to Floor 4.

State the formula that you use and show your working.

formula

working

..... J [2]

(b) (i) In the restaurant, music is being played through loudspeakers.

Explain how the sound coming from the loudspeakers reaches the people in the restaurant.

.....  
.....  
..... [2]

(ii) The amplitude of the sound waves is increased.

What effect will this have on the sounds heard by the people in the restaurant?

..... [1]

For  
Examiner's  
Use

9 Fig. 9.1 shows the main steps in a method used by a student to make salts.

In separate experiments the student reacted the carbonate of a metal and the oxide of a metal with dilute sulfuric acid.

For  
Examiner's  
Use

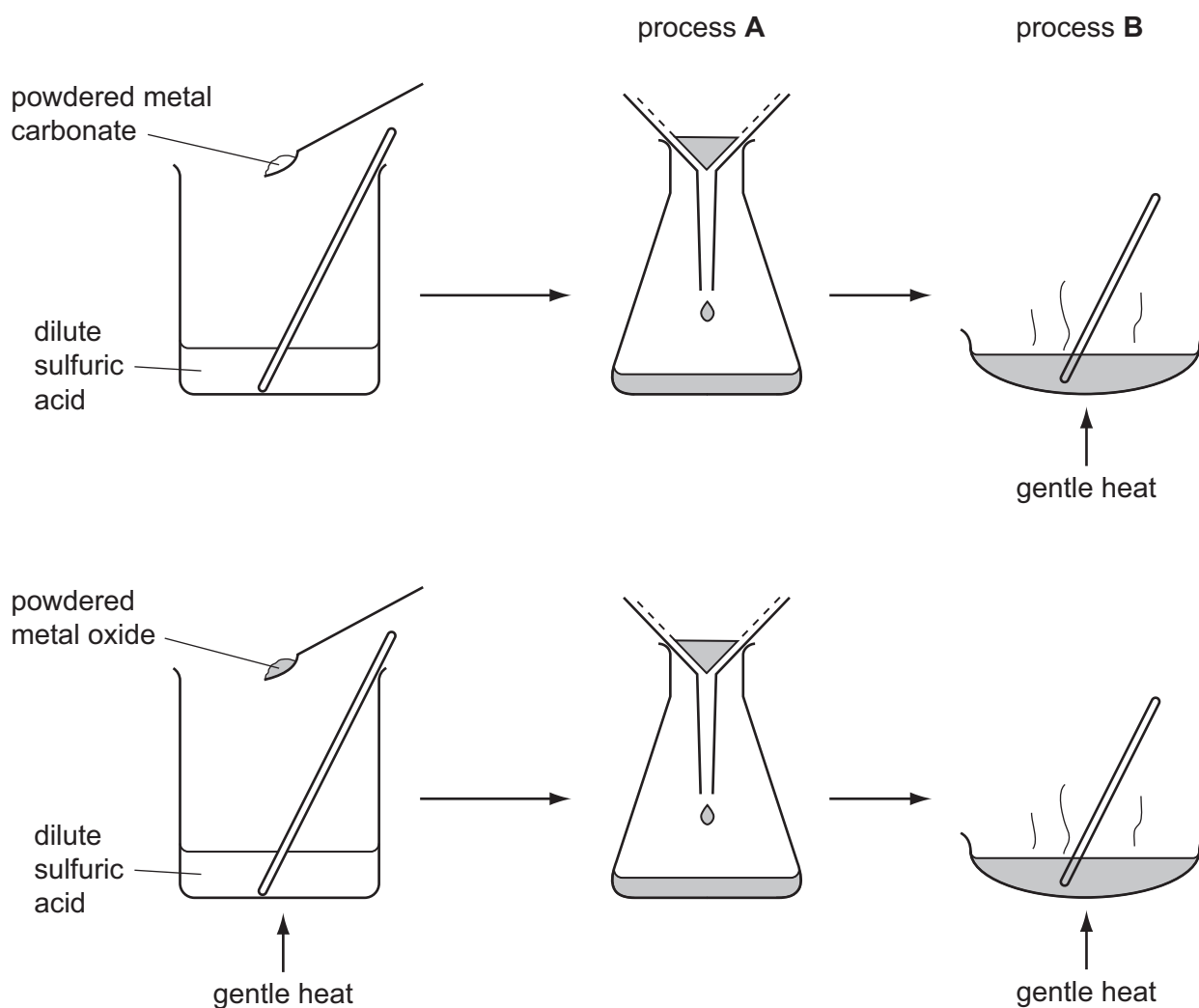


Fig. 9.1

(a) Name processes **A** and **B** shown in Fig. 9.1.

process **A** .....

process **B** ..... [2]

(b) Suggest and explain why the student used **powdered** solids in the reactions with dilute sulfuric acid.

.....

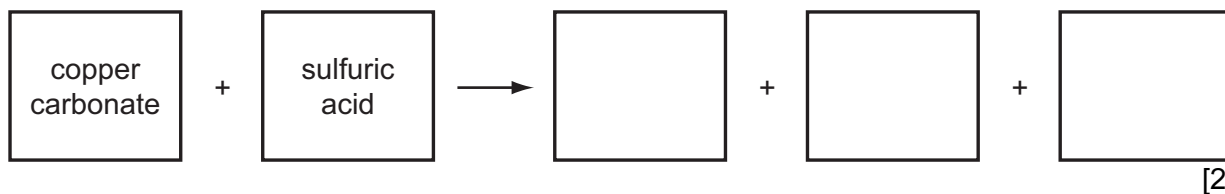
.....

..... [2]

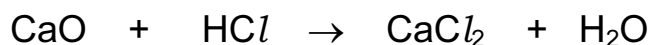
(c) (i) Name the salt which is produced when zinc oxide reacts with dilute sulfuric acid.

..... [1]

(ii) Complete the word equation for the reaction of copper carbonate with sulfuric acid.



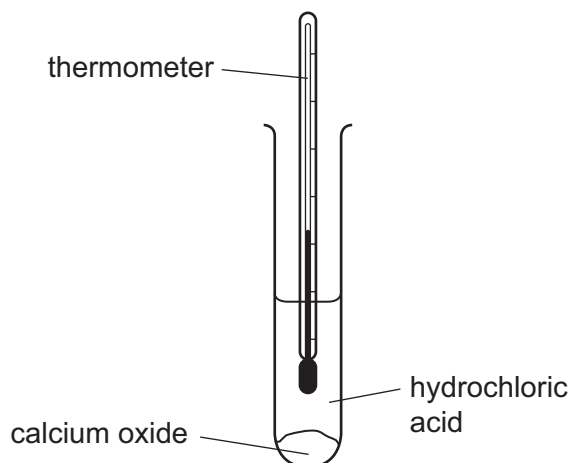
(d) (i) The salt calcium chloride is made when calcium oxide reacts with hydrochloric acid. The symbolic equation for this reaction is shown below.



Explain whether or not this equation is balanced.

.....  
 .....  
 ..... [2]

(ii) A student reacted calcium oxide with hydrochloric acid using the apparatus shown in Fig. 9.2.



**Fig. 9.2**

The student noticed that the temperature of the mixture increased.

Explain this observation.

.....  
 ..... [1]





**DATA SHEET**  
**The Periodic Table of the Elements**

		Group												
I	II	III	IV	V	VI	VII	0							
7 <b>Li</b> Lithium 3	9 <b>Be</b> Beryllium 4	1 <b>H</b> Hydrogen 1	11 <b>B</b> Boron 5	12 <b>C</b> Carbon 6	14 <b>N</b> Nitrogen 7	16 <b>O</b> Oxygen 8	19 <b>F</b> Fluorine 9	20 <b>Ne</b> Neon 10	27 <b>Al</b> Aluminium 13	28 <b>Si</b> Silicon 14	31 <b>P</b> Phosphorus 15	32 <b>S</b> Sulfur 16	35.5 <b>Cl</b> Chlorine 17	40 <b>Ar</b> Argon 18
39 <b>K</b> Potassium 19	40 <b>Ca</b> Calcium 20	55 <b>Mn</b> Manganese 25	56 <b>Fe</b> Iron 26	59 <b>Co</b> Cobalt 27	59 <b>Ni</b> Nickel 28	64 <b>Cu</b> Copper 29	65 <b>Zn</b> Zinc 30	70 <b>Ga</b> Gallium 31	73 <b>Ge</b> Germanium 32	75 <b>As</b> Arsenic 33	79 <b>Se</b> Selenium 34	80 <b>Br</b> Bromine 35	84 <b>Kr</b> Krypton 36	
85 <b>Rb</b> Rubidium 37	88 <b>Sr</b> Strontium 38	91 <b>Zr</b> Zirconium 40	91 <b>Nb</b> Niobium 41	93 <b>Nb</b> Niobium 41	96 <b>Mo</b> Molybdenum 42	106 <b>Pd</b> Palladium 46	108 <b>Ag</b> Silver 47	115 <b>In</b> Indium 49	119 <b>Sn</b> Tin 50	122 <b>Sb</b> Antimony 51	128 <b>Te</b> Tellurium 52	127 <b>I</b> Iodine 53	131 <b>Xe</b> Xenon 54	
133 <b>Cs</b> Caesium 55	137 <b>Ba</b> Barium 56	184 <b>W</b> Tungsten 74	181 <b>Ta</b> Tantalum 73	190 <b>Os</b> Osmium 76	195 <b>Pt</b> Platinum 78	197 <b>Au</b> Gold 79	201 <b>Hg</b> Mercury 80	204 <b>Tl</b> Thallium 81	207 <b>Pb</b> Lead 82	209 <b>Bi</b> Bismuth 83	210 <b>Po</b> Polonium 84	210 <b>At</b> Astatine 85	210 <b>Rn</b> Radon 86	
226 <b>Ra</b> Radium 88	227 <b>Ac</b> Actinium 89	140 <b>Ce</b> Cerium 58	141 <b>Pr</b> Praseodymium 59	144 <b>Nd</b> Neodymium 60	152 <b>Eu</b> Europium 63	157 <b>Gd</b> Gadolinium 64	159 <b>Tb</b> Terbium 65	162 <b>Dy</b> Dysprosium 66	165 <b>Ho</b> Holmium 67	167 <b>Er</b> Erbium 68	169 <b>Tm</b> Thulium 69	173 <b>Yb</b> Ytterbium 70	175 <b>Lu</b> Lutetium 71	
232 <b>Th</b> Thorium 90	238 <b>U</b> Uranium 92	232 <b>Th</b> Thorium 90	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	238 <b>U</b> Uranium 92	

\*58-71 Lanthanoid series  
†90-103 Actinoid series

a = relative atomic mass

X = atomic symbol

b = proton (atomic) number

Key

	<b>X</b>		
a	X	b	†

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

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