



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS  
International General Certificate of Secondary Education

CANDIDATE  
NAME

CENTRE  
NUMBER

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**COMBINED SCIENCE**

**0653/22**

Paper 2 (Core)

**October/November 2010**

**1 hour 15 minutes**

Candidates answer on the Question Paper.

No Additional Materials are required.

**READ THESE INSTRUCTIONS FIRST**

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use a soft pencil for any diagrams, graphs, tables or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

**DO NOT WRITE IN ANY BARCODES.**

Answer **all** questions.

A copy of the Periodic Table is printed on page 24.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use	
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10	
<b>Total</b>	

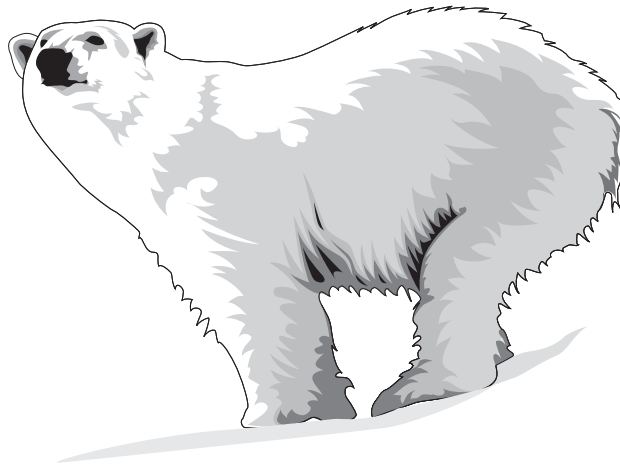
This document consists of **21** printed pages and **3** blank pages.





1 (a) Polar bears live in the cold, arctic region. They have thick, white fur.

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Describe how fur keeps a polar bear warm.

.....  
.....  
..... [2]

(b) (i) Above the arctic region the ozone layer is decreasing, allowing more ultraviolet radiation, which can cause chemical changes, to reach the surface of the Earth.

State **one** danger to human beings of being exposed to large quantities of ultraviolet radiation.

..... [1]

(ii) Ultraviolet radiation is part of the electromagnetic spectrum.

Name **one** other radiation which is part of the electromagnetic spectrum and state a use of this radiation.

name .....

use ..... [2]

- 2 (a) The apparatus shown in Fig. 2.1 can be used to react lead oxide and carbon.

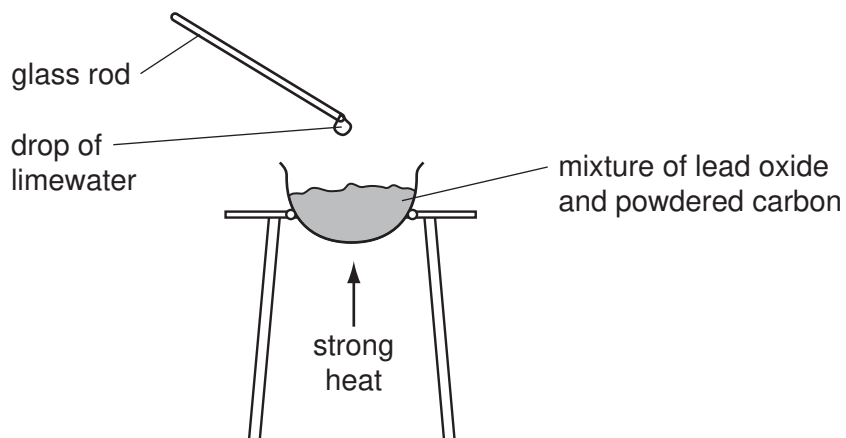


Fig. 2.1

When the mixture is heated, molten metal is formed in the container and the drop of lime water on the end of the glass rod becomes cloudy.

- (i) Suggest the **word** equation for the reaction between lead oxide and carbon. Do **not** write a symbolic equation.

..... [2]

- (ii) State **one** substance, shown in your equation in (i), which is a compound.

Explain why this substance is described as a compound and **not** as an element.

substance .....

explanation .....

..... [3]

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- (b) Fig. 2.2 shows some of the apparatus used in the electrolysis of copper chloride solution.

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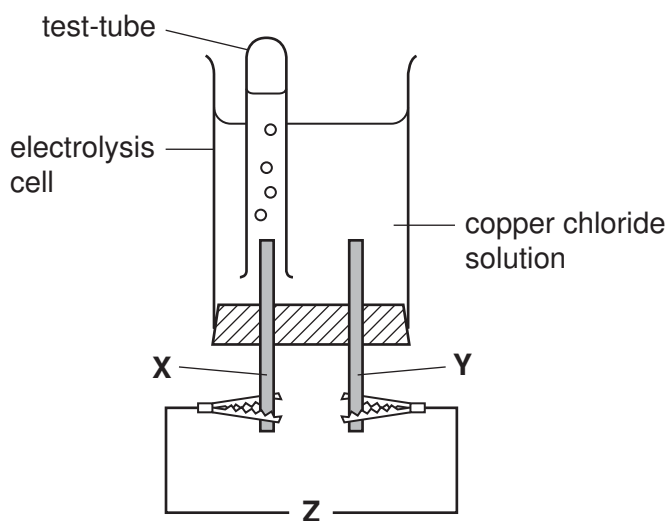


Fig. 2.2

- (i) What is missing from position **Z** in Fig. 2.2?

..... [1]

- (ii) Name the gas which collects in the test-tube, and explain whether electrode **X** is the anode or the cathode.

gas .....

Electrode **X** is the ..... because .....

.....  
..... [2]

3 A healthy plant growing in a pot was watered and placed in a sunny window. A transparent plastic bag was placed over the plant, as shown in Fig. 3.1.

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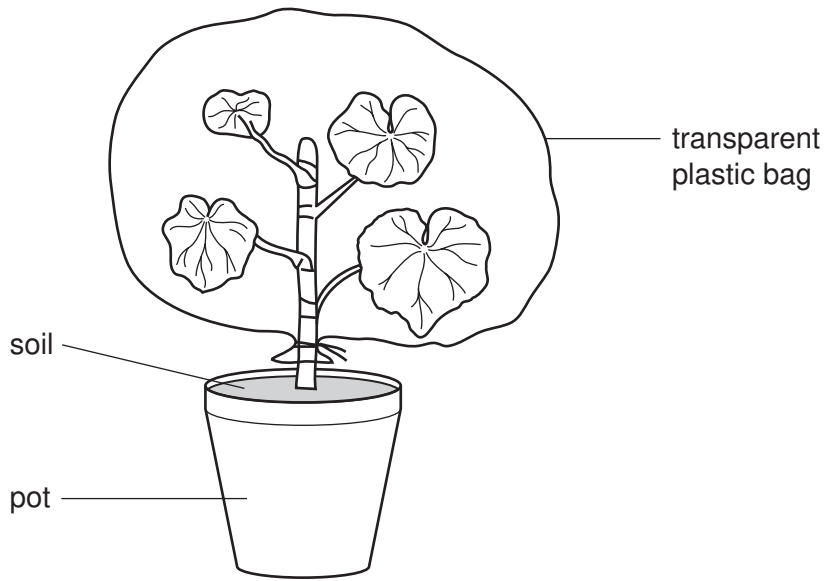


Fig. 3.1

(a) The temperature near the window fell overnight. The next morning, small droplets of liquid water were visible on the inside of the plastic bag.

(i) Name the process by which plant leaves lose water vapour.

..... [1]

(ii) Name the small holes in the leaf through which the water vapour is lost.

..... [1]

(iii) Explain why the water formed droplets of liquid on the plastic bag.

.....  
.....  
..... [2]

(b) Fig. 3.2 shows a cell from the plant leaf.

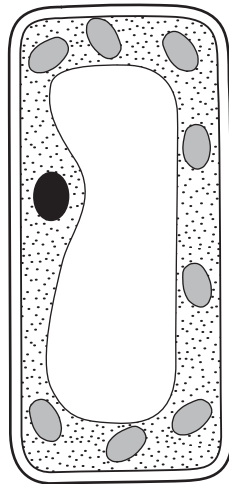


Fig. 3.2

(i) On the diagram of the cell in Fig. 3.2, label and name **two** structures that would **not** be present in an animal cell. [2]

(ii) Name the part of the leaf in which this cell could be found.

..... [1]

(iii) The cell in Fig. 3.2 can photosynthesise.

Write the word equation for photosynthesis.



[2]

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- 4 (a) Fig. 4.1 shows the speed-time graph for a train.

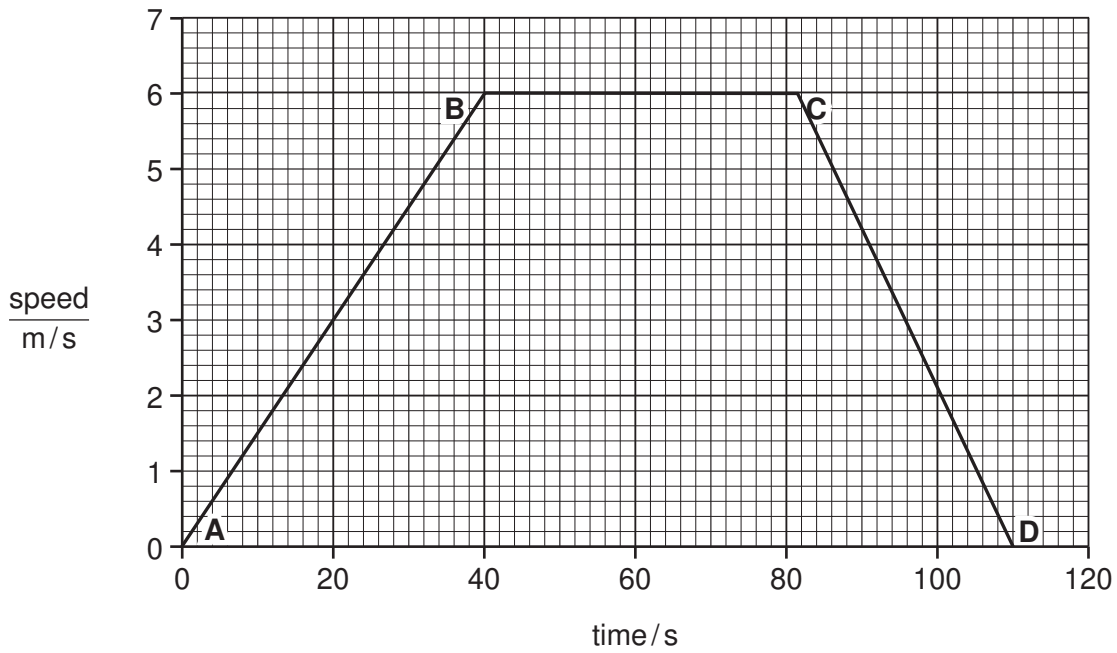


Fig. 4.1

The brakes are applied at C. Calculate how long it takes the train to stop.

.....s [1]

- (b) Another train, on a journey lasting 10 minutes, travelled at a constant speed of 9 m/s.

- (i) Show that the distance travelled by the train during this journey was 5400 m.

State the formula that you use and show your working.

formula used

working

[2]



- (ii) The average force needed for the train to maintain the speed of 9 m/s was 10 000 N.

Calculate the work done by the train over 10 minutes.

State the formula that you use and show your working.

formula used

working

..... J [2]

For  
Examiner's  
Use

5 Fig. 5.1 shows some stages in the formation of a human fetus.

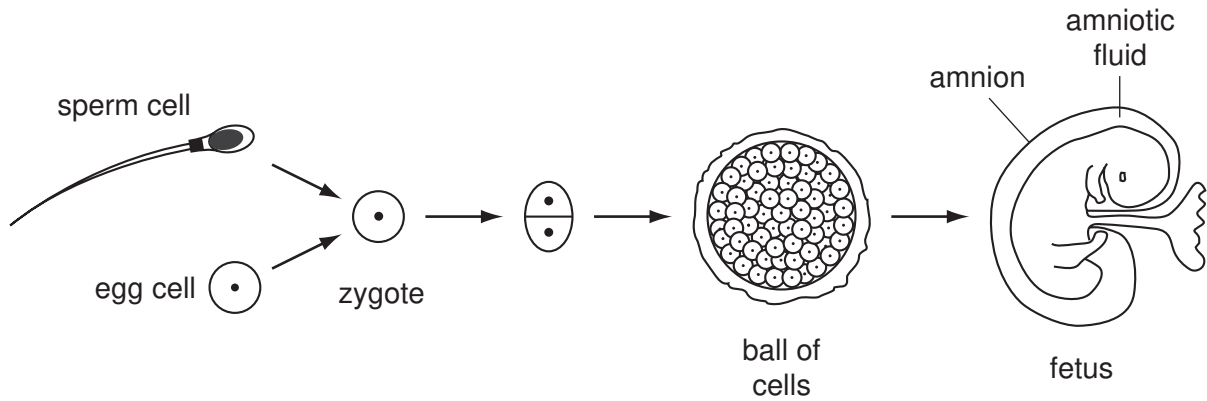


Fig. 5.1

For  
Examiner's  
Use

(a) Most human cells contain 46 chromosomes, but egg cells and sperm cells contain only 23 chromosomes each.

Suggest a reason for this.

.....  
 ..... [1]

(b) Name the part of the reproductive system in which each of these events takes place.

(i) Eggs are produced. .... [1]

(ii) Fertilisation. .... [1]

(c) Describe the function of the amnion.

.....  
 .....  
 ..... [2]

(d) The fetus develops in the uterus.

It is attached to the uterus by the umbilical cord and placenta.

It obtains nutrients from its mother's blood, through the placenta.

Suggest why a pregnant woman should have more iron and calcium in her diet than when she is not pregnant.

iron .....

.....

calcium .....

..... [3]

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- 6 (a) Electrical equipment can be dangerous, especially when it is handled with wet hands.

Explain why you are quite likely to be electrocuted if you handle an electrical device with wet hands rather than dry hands.

.....  
 ..... [1]

- (b) Fig. 6.1 shows a simple electric circuit.

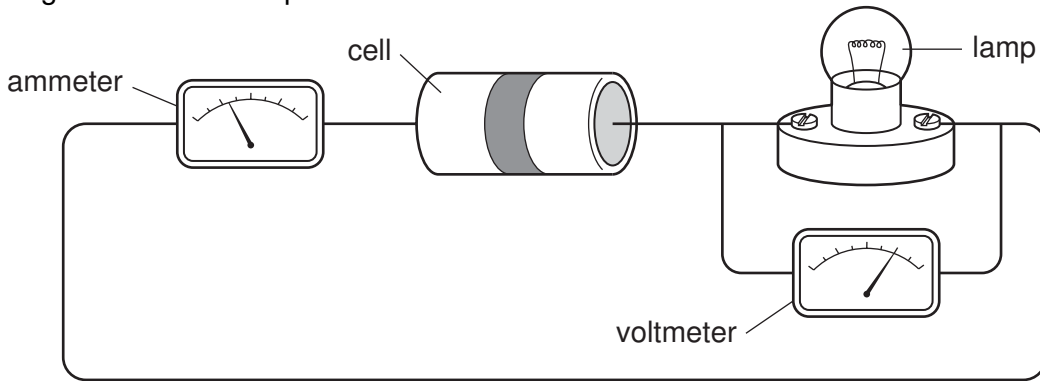


Fig. 6.1

Draw the circuit diagram for the circuit in Fig. 6.1 using the correct symbols.

[3]

(c) Fig. 6.2 shows a circuit built by a student.

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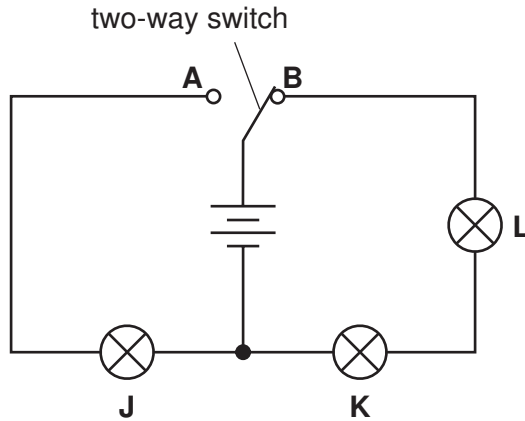


Fig. 6.2

(i) The switch is at position **B**.

Which lamps will be lit? .....

[1]

(ii) The switch is then moved to position **A**.

What happens to lamps **J**, **K** and **L**?

lamp **J** .....

lamp **K** .....

lamp **L** .....

[2]

(d) The student has six resistors as shown in Fig. 6.3.

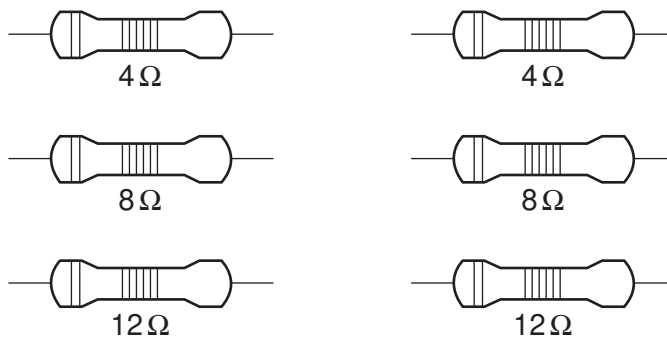


Fig. 6.3

Describe how he can combine **two** of these resistors to get a total resistance of 20 ohms.

.....  
.....

[1]

(e) Power stations produce electricity.

Six stages in the production of electricity at a coal-fired power station are shown below.

- A electricity produced
- B coal burned
- C steam produced
- D turbine driven by steam
- E turbine turns generator
- F water boils

Using the letters **A** to **F**, list the stages in the correct order in the boxes below. Two have been done for you.

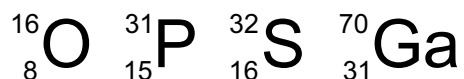


[2]

For  
Examiner's  
Use

Please turn over for Question 7.

- 7 (a) The chemical symbols for the atoms shown below include proton (atomic) numbers and nucleon (mass) numbers.



- (i) State which of these symbols represent atoms of elements in the same **group** of the Periodic Table

..... [1]

- (ii) Complete Table 7.1 which shows the names and the numbers of protons and neutrons in two of the atoms shown above.

Table 7.1

element name	protons	neutrons
oxygen		
	15	16

[2]

- (b) Chlorine and hydrogen combine to form hydrogen chloride which dissolves in water to produce hydrochloric acid.

- (i) Suggest a substance which reacts with hydrochloric acid to form the salt, copper chloride.

..... [1]

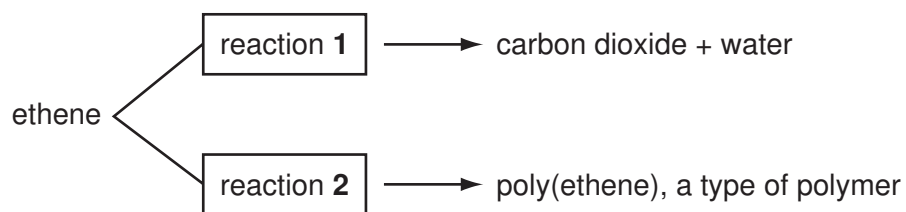
- (ii) Suggest an element from the third period of the Periodic Table which reacts **safely** with hydrochloric acid to produce hydrogen gas.

..... [1]



(c) Ethene is a gaseous compound of carbon and hydrogen.

Fig. 7.2 shows two different chemical reactions, **1** and **2**, involving ethene.



**Fig. 7.2**

(i) For reactions **1** and **2**, deduce the type of chemical reaction which occurs.

reaction **1** .....

reaction **2** ..... [2]

(ii) For reaction **2**, describe briefly what happens to the molecules of ethene during the reaction.

.....

..... [1]

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8 Soya beans are an important crop in many tropical and subtropical countries, because they contain a lot of protein.

For  
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Use

(a) Fig. 8.1 shows how the yield of soya beans is affected by the pH of the soil in which they are grown.

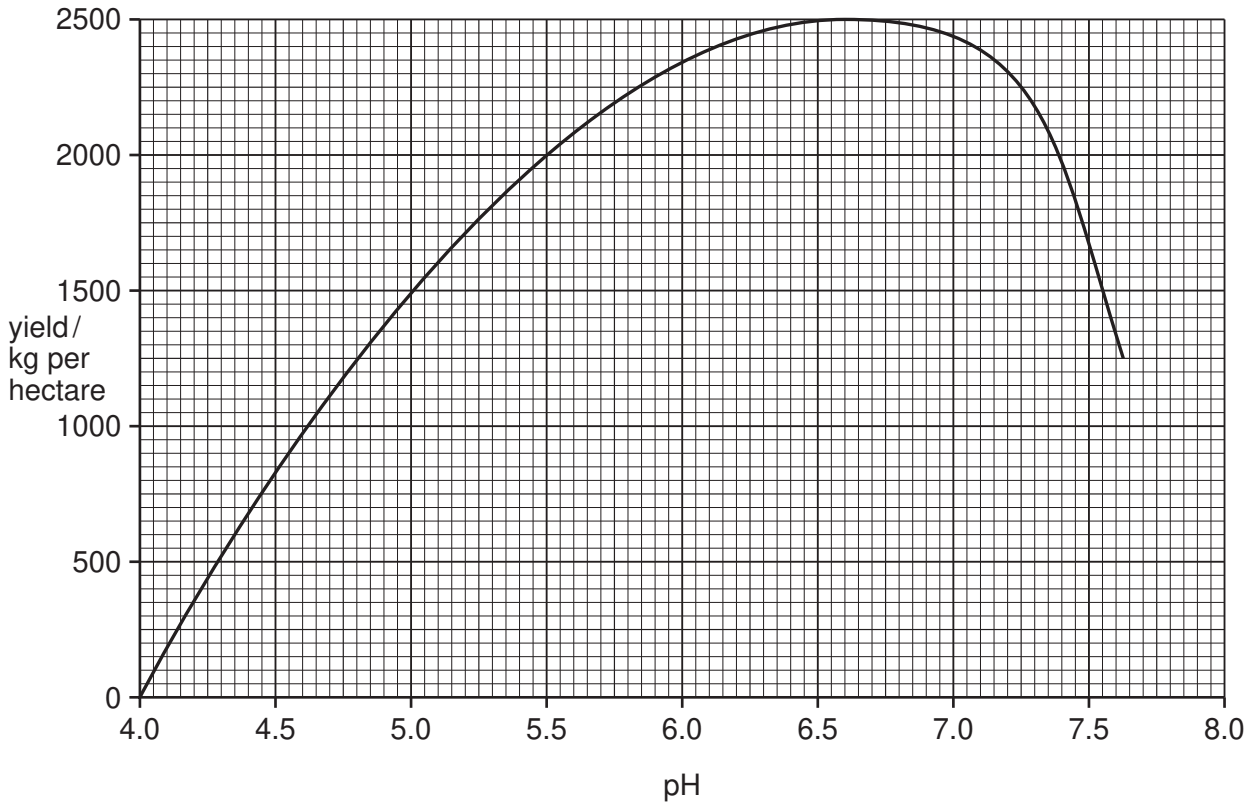


Fig. 8.1

A farmer grows soya beans in a field where the soil has a pH of 5.5.

(i) What yield of beans could he get from his crop?

..... kg per hectare [1]

(ii) State the pH range in which soya beans grow best.

between ..... and ..... [1]

(iii) The farmer decides to add calcium carbonate to the soil in his field.

Explain why this would help him to achieve a higher yield of soya beans.

.....  
 .....  
 ..... [2]

(b) The field is on a steep slope.

Describe **two** things the farmer could do to reduce the risk of soil erosion.

1 .....

.....

2 .....

..... [2]

(c) Soya beans are seeds. They grow after the flowers on the soya plants have been pollinated.

(i) Soya flowers often have violet-coloured petals.

Suggest how soya flowers are pollinated.

..... [1]

(ii) Explain why soya beans only grow after the flowers have been pollinated.

.....

.....

..... [2]

(iii) Describe how you would test a soya bean seed for protein. State the result you would expect.

test .....

.....

result ..... [2]

- 9 (a) Complete Table 9.1 to show the properties of alpha, beta and gamma radiations.

For  
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Table 9.1

	description	charge	range in air	ionising ability
alpha		positive	5 cm	very strong
beta	electron		50 cm	
gamma	wave		many kilometres	weak

[4]

- (b) Many people have smoke detectors in their houses.

Smoke detectors contain a radioactive source which emits alpha radiation.

Explain why the alpha radiation from the smoke detector is not dangerous to people living in the house.

.....

.....

..... [1]

10 In many countries, river water is collected and treated to make it safe for humans to drink.

(a) State and explain which **two** of the processes shown below are used to treat river water so that it becomes safe to drink.

adding chlorine                  chromatography                  evaporation                  filtration

first process .....

explanation

.....

second process .....

explanation .....

..... [4]

(b) Sulfur dioxide is a gaseous compound which is released into the air when fossil fuels containing sulfur compounds are burned.

(i) Describe how sulfur dioxide gas could cause pollution of water in rivers and lakes.

.....

.....

.....

..... [3]

(ii) Suggest **one** way in which sulfur dioxide emissions into the atmosphere are being reduced.

.....

..... [1]

(c) Fig. 10.1 shows a diagram of a water molecule, H<sub>2</sub>O.

Choose words or phrases from the following list to complete the labelling of the diagram.

**covalent bond**

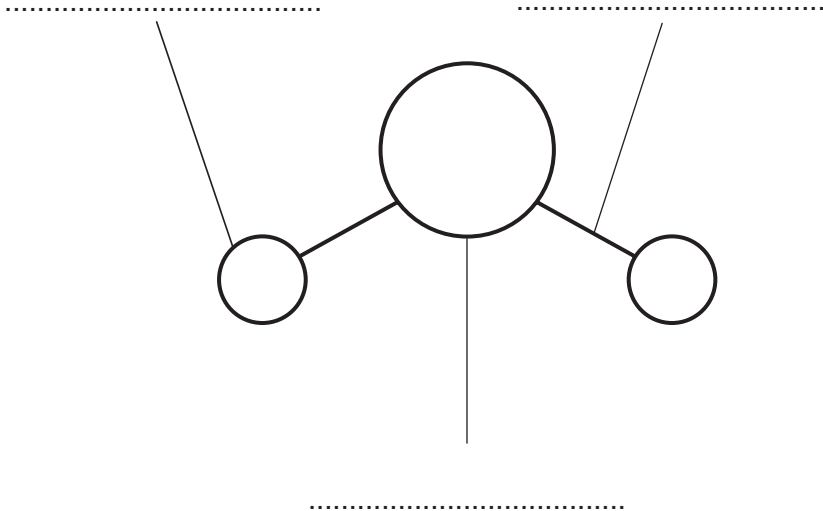
**hydrogen atom**

**ionic bond**

**nucleus**

**oxygen atom**

**proton**



**Fig. 10.1**

[3]

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Use*



**DATA SHEET**  
**The Periodic Table of the Elements**

		Group																					
		I	II	III	IV	V	VI	VII	0														
		<table border="1" style="width: 100%; border-collapse: collapse;"> <tr> <td style="width: 10%; text-align: center;">1</td> <td style="width: 10%; text-align: center;"><b>H</b> Hydrogen 1</td> <td colspan="8"></td> <td style="width: 10%; text-align: center;">2</td> <td style="width: 10%; text-align: center;"><b>He</b> Helium 2</td> </tr> </table>										1	<b>H</b> Hydrogen 1									2	<b>He</b> Helium 2
1	<b>H</b> Hydrogen 1									2	<b>He</b> Helium 2												
3	7	9											10	20									
<b>Li</b> Lithium 3	<b>Be</b> Beryllium 4	<b>B</b> Boron 5	<b>C</b> Carbon 6	<b>N</b> Nitrogen 7	<b>O</b> Oxygen 8	<b>F</b> Fluorine 9	<b>Ne</b> Neon 10						<b>Ar</b> Argon 18										
11	23	24											16	32									
<b>Na</b> Sodium 11	<b>Mg</b> Magnesium 12	<b>Al</b> Aluminium 13	<b>Si</b> Silicon 14	<b>P</b> Phosphorus 15	<b>S</b> Sulfur 16	<b>Cl</b> Chlorine 17	<b>Ar</b> Argon 18						<b>Kr</b> Krypton 36										
19	39	40											34	70									
<b>K</b> Potassium 19	<b>Ca</b> Calcium 20	<b>Sc</b> Scandium 21	<b>Ti</b> Titanium 22	<b>V</b> Vanadium 23	<b>Cr</b> Chromium 24	<b>Mn</b> Manganese 25	<b>Fe</b> Iron 26	<b>Co</b> Cobalt 27	<b>Ni</b> Nickel 28	<b>Cu</b> Copper 29	<b>Zn</b> Zinc 30	<b>Ga</b> Gallium 31	<b>Ge</b> Germanium 32	<b>As</b> Arsenic 33	<b>Se</b> Selenium 34	<b>Br</b> Bromine 35	<b>Kr</b> Krypton 36						
37	85	88											81	112									
<b>Rb</b> Rubidium 37	<b>Sr</b> Strontium 38	<b>Y</b> Yttrium 39	<b>Zr</b> Zirconium 40	<b>Nb</b> Niobium 41	<b>Mo</b> Molybdenum 42	<b>Tc</b> Technetium 43	<b>Ru</b> Ruthenium 44	<b>Rh</b> Rhodium 45	<b>Pd</b> Palladium 46	<b>Ag</b> Silver 47	<b>Cd</b> Cadmium 48	<b>In</b> Indium 49	<b>Sn</b> Tin 50	<b>Sb</b> Antimony 51	<b>Te</b> Tellurium 52	<b>I</b> Iodine 53	<b>Xe</b> Xenon 54						
55	133	137											84	207									
<b>Cs</b> Caesium 55	<b>Ba</b> Barium 56	<b>La</b> Lanthanum 57	<b>Hf</b> Hafnium 72	<b>Ta</b> Tantalum 73	<b>W</b> Tungsten 74	<b>Re</b> Rhenium 75	<b>Os</b> Osmium 76	<b>Ir</b> Iridium 77	<b>Pt</b> Platinum 78	<b>Au</b> Gold 79	<b>Hg</b> Mercury 80	<b>Tl</b> Thallium 81	<b>Pb</b> Lead 82	<b>Bi</b> Bismuth 83	<b>Po</b> Polonium 84	<b>At</b> Astatine 85	<b>Rn</b> Radon 86						
87	226	227											88	209									
<b>Fr</b> Francium 87	<b>Ra</b> Radium 88	<b>Ac</b> Actinium 89											89	209									

140	<b>Ce</b> Cerium 58	141	<b>Pr</b> Praseodymium 59	144	<b>Nd</b> Neodymium 60	150	<b>Sm</b> Samarium 62	152	<b>Eu</b> Europium 63	157	<b>Gd</b> Gadolinium 64	162	<b>Dy</b> Dysprosium 66	165	<b>Ho</b> Holmium 67	167	<b>Er</b> Erbium 68	169	<b>Tm</b> Thulium 69	173	<b>Yb</b> Ytterbium 70	175	<b>Lu</b> Lutetium 71				
90	<b>Th</b> Thorium 90	91	<b>Pa</b> Protactinium 91	92	<b>U</b> Uranium 92	93	<b>Np</b> Neptunium 93	94	<b>Pu</b> Plutonium 94	95	<b>Am</b> Americium 95	96	<b>Cm</b> Curium 96	97	<b>Bk</b> Berkelium 97	98	<b>Cf</b> Californium 98	99	<b>Es</b> Einsteinium 99	100	<b>Fm</b> Fermium 100	101	<b>Md</b> Mendelevium 101	102	<b>No</b> Nobelium 102	103	<b>Lr</b> Lawrencium 103

a	<b>X</b>	b
Key	a = relative atomic mass	X = atomic symbol
		b = proton (atomic) number

\*58-71 Lanthanoid series  
†90-103 Actinoid series

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

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