



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

www.XtremePapers.com

COMBINED SCIENCE

0653/13

Paper 1 Multiple Choice

October/November 2011

45 minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)



READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

This document consists of **14** printed pages and **2** blank pages.

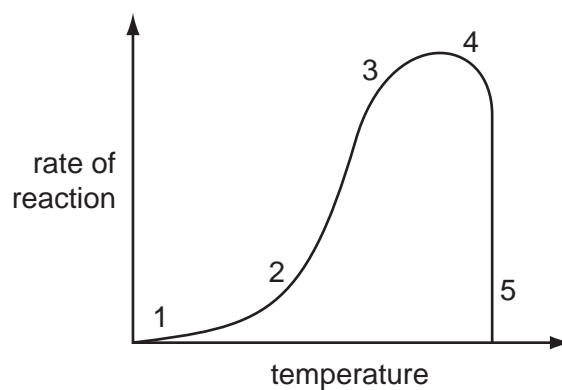


- 1 Which part of a plant cell is made of cellulose?
- A cell membrane
 B cell wall
 C chloroplast
 D nucleus
- 2 Which leaf tissue has specialised cells that surround stomata?
- A epidermis
 B palisade mesophyll
 C phloem
 D xylem

- 3 Which parts of a cell control its activities and control what enters and leaves it?

	controls cell's activities	controls what enters and leaves the cell
A	chloroplast	cell surface membrane
B	chloroplast	cell wall
C	nucleus	cell surface membrane
D	nucleus	cell wall

- 4 The graph shows the effect of temperature on the rate of an enzyme-controlled reaction.



Where on the graph has all the enzyme been denatured?

- A 1 B 2 and 3 C 3 and 4 D 5

- 5 Oxygenated blood returns to the heart from the lungs in vessel X and leaves the heart to circulate around the body in vessel Y.

What are X and Y?

	X	Y
A	aorta	pulmonary vein
B	pulmonary artery	vena cava
C	pulmonary vein	aorta
D	vena cava	pulmonary artery

- 6 When a leaf is photosynthesising, in which direction do gases diffuse through the stomata?

	carbon dioxide	oxygen
A	in	in
B	in	out
C	out	in
D	out	out

- 7 What happens during digestion?

	large pieces of food are broken into small pieces	large molecules are broken into small molecules
A	✓	✓
B	✓	x
C	x	✓
D	x	x

- 8 A species of animal reproduces both sexually and asexually.

Which offspring will be clones?

	offspring from sexual reproduction	offspring from asexual reproduction
A	✓	✓
B	✓	x
C	x	✓
D	x	x

- 9 The table shows the level of alcohol in a person's blood after drinking two litres of beer.

time after drinking beer (hours)	alcohol in the blood (grams/dm ³)
1	7
2	5
3	3
4	0

How long will it be (in hours) before the person's reaction time returns to normal?

- A 0 to 1 B 1 to 2 C 2 to 3 D 3 to 4
- 10 Which method of family planning is also likely to reduce the risk of the spread of syphilis?
- A condom
 B intra-uterine device (IUD)
 C pill
 D sterilisation
- 11 Which process reduces soil erosion on hilly ground?
- A cutting down the trees
 B increasing the number of grazing animals
 C ploughing up and down the hilly ground
 D terracing the hilly ground
- 12 Albino humans cannot make any pigment in their skin.

A pale-skinned student, who is **not** an albino, sits in the sun on a number of days. The student's skin becomes suntanned (darker).

What causes this suntanning to happen?

- A the environment and the student's albino alleles
 B the environment and the student's non-albino alleles
 C the environment only
 D the student's genes only

13 The diagram shows a food chain.



Which types of energy are represented by the black arrows and by the white arrows?

	black arrows	white arrows
A	chemical	heat
B	chemical	light
C	heat	chemical
D	light	chemical

14 Element X has a nucleon number of 40.

The electron arrangement of element X is 2,8,8.

Which statements about element X are correct?

- 1 It has 40 neutrons in its nucleus.
- 2 It has 2 electrons in its outer shell.
- 3 It is unreactive.
- 4 It is in Group 0 of the Periodic Table.

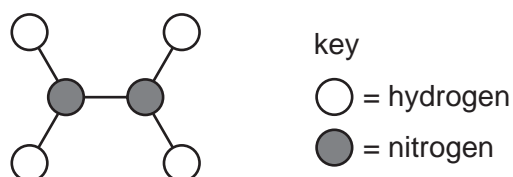
A 1 and 2

B 1 and 3

C 2 and 4

D 3 and 4

15 A model of a molecule is shown.

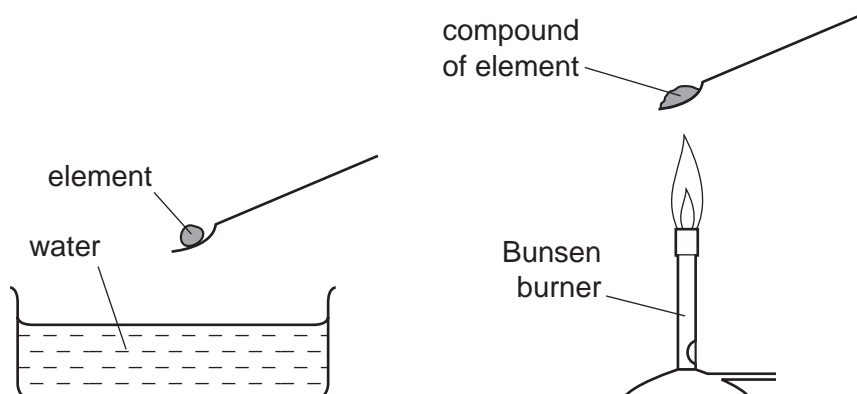


Which description and formula are correct for this molecule?

	description	formula
A	compound	NH ₂
B	compound	N ₂ H ₄
C	mixture	NH ₂
D	mixture	N ₂ H ₄

- 16 In an experiment the elements calcium, copper, potassium and sodium were separately reacted with water.

In a second experiment a flame test was carried out on compounds of each of the elements.



Which row correctly shows the reaction of the elements with water and the colour of the flame?

	element	reaction with water	colour of the flame
A	calcium	vigorous	green
B	copper	no reaction	red
C	potassium	vigorous	lilac
D	sodium	no reaction	yellow

- 17 Which two elements do **not** form an alloy?

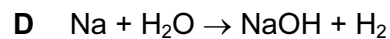
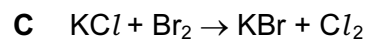
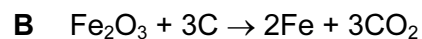
- A** carbon and sulfur
- B** carbon and iron
- C** copper and zinc
- D** silver and gold

- 18 Sulfur dioxide is formed as a pollutant when fossil fuels are burned.

Which properties does sulfur dioxide have?

	toxic	acidic	corrosive
A	✓	✓	✓
B	✓	✓	x
C	✓	x	x
D	x	x	x

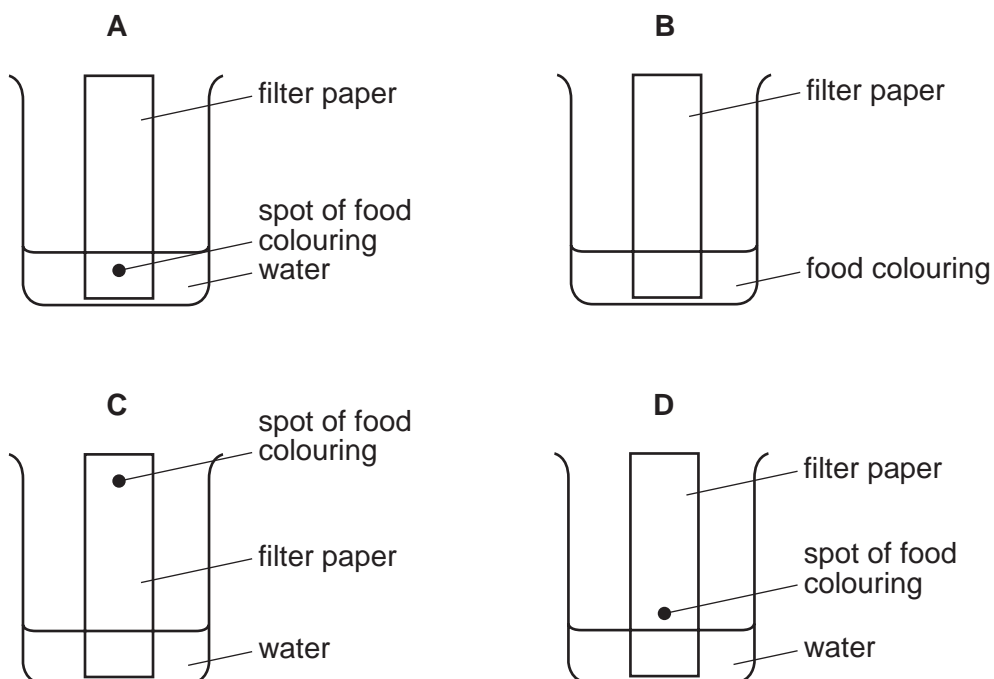
19 Which equation is correctly balanced?



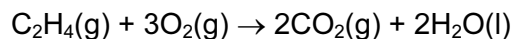
20 A soft metal reacts vigorously with cold water.

Which letter shows the position of this metal in the Periodic Table?

21 Which diagram shows how a mixture of dyes in a food colouring are separated?



22 Ethene burns as shown.

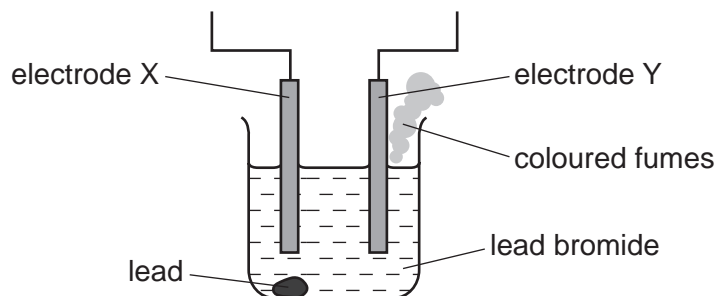


What happens to ethene in this reaction?

- A decomposition
- B neutralisation
- C oxidation
- D reduction

23 The diagram shows the electrolysis of lead(II) bromide using inert electrodes.

Lead is formed at electrode X and coloured fumes at electrode Y.



Which statement about the electrolysis of lead(II) bromide is correct?

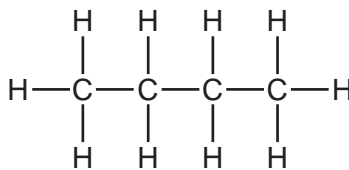
- A Electrode X is the anode.
- B The colour of the fumes is brown.
- C The lead(II) bromide is in aqueous solution.
- D The mass of the lead(II) bromide does not change during the reaction.

24 When compound X is added to pure water, the pH increases.

Which formula could **not** be a correct formula for X?

- A HNO_3
- B KOH
- C NaOH
- D NH_3

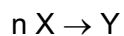
25 The structure of a molecule is shown.



Which term correctly describes this molecule?

- A hydrocarbon
- B monomer
- C petroleum
- D polymer

26 Many molecules of X combine to form a single molecule Y as shown in the equation.



(n is a very large number)

Which terms best describe X and Y in this reaction?

	X	Y
A	fraction	monomer
B	monomer	fraction
C	monomer	polymer
D	polymer	fraction

27 Which change does **not** alter the rate of reaction between zinc and dilute sulfuric acid?

- A addition of a catalyst
- B change in concentration of the acid
- C change in atmospheric pressure
- D change in temperature

28 What is the meaning of the *weight* of an object?

- A the density of the material from which it is made
- B the force exerted on it by gravity
- C the mass of the matter it contains
- D the pressure it exerts on the ground

29 The table gives information about a liquid in a container.

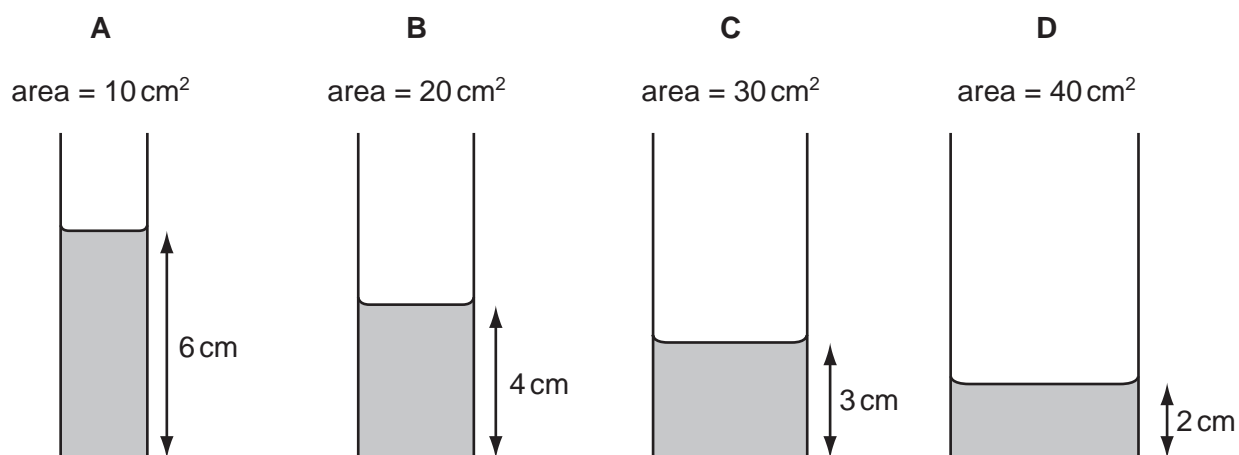
depth of liquid	10 cm
mass of liquid	30 g
temperature of liquid	25 °C
volume of liquid	20 cm ³

What is the density of the liquid?

- A** 0.33 cm/g **B** 1.2 g/°C **C** 1.5 g/cm³ **D** 3.0 g/cm

30 Some water is poured into four tubes of different cross-sectional areas.

Which tube holds the largest volume of water?



31 An object travels 6.0 km in 2 minutes.

What is its speed?

- A** 0.050 m/s **B** 3.0 m/s **C** 50 m/s **D** 3000 m/s

32 Which source releases energy by burning when it is used in the process of generating electricity?

- A** a fossil fuel
B hydroelectric
C nuclear
D solar

33 When flying, some birds use warm air currents to gain height.

What is the cause of these currents?

- A conduction
- B convection
- C evaporation
- D radiation

34 Why is a fuse used in an electric circuit in a house?

- A to increase the resistance of the circuit
- B to keep the power used to a minimum value
- C to prevent a short circuit from occurring
- D to stop the cables overheating

35 Diagram 1 shows two identical resistors R_1 and R_2 connected in series in a circuit.

R_2 is then removed, as shown in diagram 2.

diagram 1

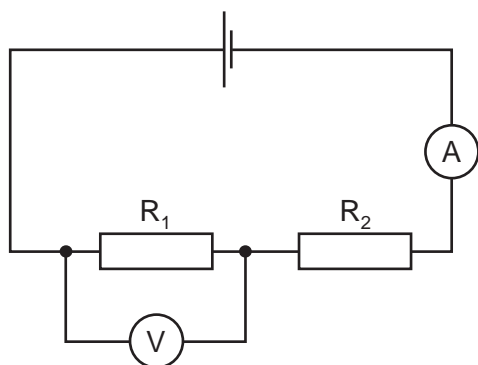
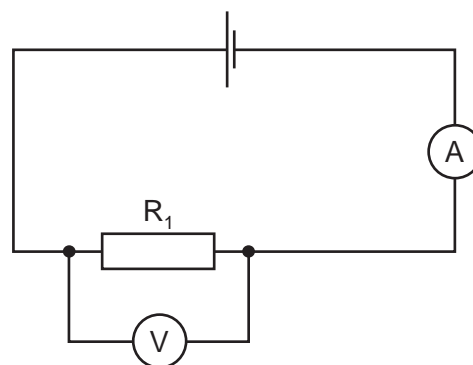


diagram 2



How do the readings on the ammeter and the voltmeter change when R_2 is removed?

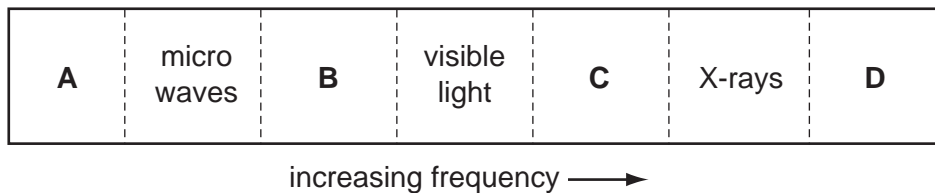
	ammeter	voltmeter
A	decreases	decreases
B	decreases	increases
C	increases	decreases
D	increases	increases

36 Which row shows two of the essential items used in the construction of a transformer?

	iron core	permanent magnet	primary coil	slip rings
A	✓	✓		
B	✓		✓	
C		✓		✓
D			✓	✓

37 The diagram shows the spectrum of electromagnetic waves.

Which labelled region represents gamma rays?

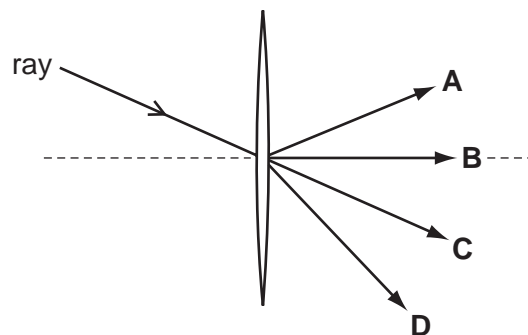


38 Which is the best description of a wave that is a quiet, high-pitched sound?

- A** large amplitude and high frequency.
- B** large amplitude and low frequency.
- C** small amplitude and high frequency.
- D** small amplitude and low frequency.

39 A ray of light passes through the centre of a thin converging lens.

In which direction does the ray leave the lens?



40 Which nuclear process occurs in the Sun, and which process is used in a nuclear power station?

	in the Sun	in a nuclear power station
A	fission	fission
B	fission	fusion
C	fusion	fission
D	fusion	fusion

DATA SHEET
The Periodic Table of the Elements

		Group												
		I	II	III	IV	V	VI	VII	VIII	IX	X			
		1 H Hydrogen 1												
7	9	3	4	5	6	7	8	9	10	11	12			
Li Lithium	Be Beryllium	B Boron	C Carbon	N Nitrogen	O Oxygen	F Fluorine	Ne Neon	Na Sodium	Mg Magnesium	Al Aluminium	Si Silicon			
11	12	13	14	15	16	17	18	19	20	21	22			
Na Sodium	Mg Magnesium	Al Aluminium	Si Silicon	P Phosphorus	S Sulfur	Cl Chlorine	Ar Argon	K Potassium	Ca Calcium	Sc Scandium	Ti Titanium			
19	20	21	22	23	24	25	26	27	28	29	30			
K Potassium	Ca Calcium	Sc Scandium	Ti Titanium	V Vanadium	Cr Chromium	Mn Manganese	Fe Iron	Co Cobalt	Ni Nickel	Cu Copper	Zn Zinc			
37	38	39	40	41	42	43	44	45	46	47	48			
Rb Rubidium	Sr Strontium	Y Yttrium	Zr Zirconium	Nb Niobium	Mo Molybdenum	Tc Technetium	Ru Ruthenium	Rh Rhodium	Pd Palladium	Ag Silver	Cd Cadmium			
55	56	57	72	73	74	75	76	77	78	79	80			
Cs Caesium	Ba Barium	La Lanthanum	Hf Hafnium	Ta Tantalum	W Tungsten	Re Rhenium	Os Osmium	Ir Iridium	Pt Platinum	Au Gold	Hg Mercury			
87	88	89	†	†	†	†	†	†	†	†	†			
Fr Francium	Ra Radium	Ac Actinium												
<p>*58-71 Lanthanoid series †90-103 Actinoid series</p>														
<p>Key</p> <table style="margin-left: auto; margin-right: auto;"> <tr> <td style="border: 1px solid black; padding: 2px;">a</td> <td style="border: 1px solid black; padding: 2px;">X</td> <td style="border: 1px solid black; padding: 2px;">b</td> </tr> </table> <p>a = relative atomic mass X = atomic symbol b = proton (atomic) number</p>												a	X	b
a	X	b												

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

Permission to reproduce items where third-party owned material protected by copyright is included has been sought and cleared where possible. Every reasonable effort has been made by the publisher (UCLES) to trace copyright holders, but if any items requiring clearance have unwittingly been included, the publisher will be pleased to make amends at the earliest possible opportunity.

University of Cambridge International Examinations is part of the Cambridge Assessment Group. Cambridge Assessment is the brand name of University of Cambridge Local Examinations Syndicate (UCLES), which is itself a department of the University of Cambridge.